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None of these. Answer: A
Learning Objective: 2.2
Demonstrate how to read and

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draw bond-line structures through converting other styles of molecular representation into bond-line structures and vice versa Difficulty: Easy. 35. Which of the following is the correct bond-line structure for (CH.

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hybridization of the central carbon in $\text{CH}_3\text{C}\equiv\text{N}$ and the bond angle CCN are a. sp^2 , 180° . b. sp , 180° . c. sp^2 , 120° . d. sp^3 , 109° .

2. Which of the following statements about an sp hybridized carbon is FALSE? a. It is divalent. b. It forms bonds that are linear. c. It has two p orbitals. d.

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details. Thanks. Klein, Organic Chemistry 3e Chapter 5 1. What is the relationship between the following two compounds? A. constitutional isomers B. stereoisomers C. identical D. not isomers; different compounds entirely Answer: A Learning ...

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