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1. A solution is a homogeneous mixture of two or 9. more chemically non-reacting substances. The components of a solution generally cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

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Solution Chemistry Notes. 2. Solutions • Basic definitions – Solution: homogenous mixture – Solute: substance that is dissolved – Solvent: substance that contains the solute in a solution • Process of Dissolving: – Molecules of the solvent “pull apart” the solute molecules, and distributes them within the solvent • How to tell the solute from the solvent – The substance that is in the greater amount is the solvent – The substance that keeps its phase is the solvent.

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CBSE Class 12 Chemistry Notes: Solutions – Basics. by Anand Meena. April 22, 2019. in 12th Class, Class Notes. Reading Time: 2min read 0. 4. SHARES. 0. VIEWS. Share on Facebook Share on Twitter. Solution is a homogeneous mixture of two or more chemically non reacting substances whose composition can be varied within certain limits. A solution ...

CBSE Class 12 Chemistry Notes: Solutions - Basics ...
Solutions Class 12 Notes Chapter 2. A solution comprises a solute and a solvent. It is defined as a homogeneous mixture of two or more than

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two substances. They can be classified into three types: Solid solutions. Gaseous solutions. Liquid solutions. Molarity, mole fraction, percentages, and molality are the terms used to express the concentration of a solution.

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Revision Notes on Solution: A solution is a homogeneous mixture of two (or more) substances, the composition of which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute. If both components are in same physical state.

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defined in terms of concentration.

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NCERT Notes For Class 12 Chemistry Chapter 2: Solutions. Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent

“The Objective of Education is to prepare the young to Educate themselves throughout their Lives” This philosophy has always been followed by Gujarat Secondary Education Board (GSEB), whether through their education system framework or regular enhancement in

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curriculum. GSEB ensures better access, equality and quality in elementary education for school students. In order to achieve aforesaid objectives, Gujarat State Board of School Textbooks (GSSTB) has proposed a new syllabus for school textbooks, which will be aligned with NCERT. We at Oswaal Books, welcome the above decision of GSEB and have ensured our offerings include updated content, aligned with the latest syllabus as directed by the Board. Oswaal GSEB NCERT Solutions are designed as per the latest curriculum of Gujarat Board and emphasize on nurturing individuality thus enhancing one's innate potentials which help in increasing self-confidence. We believe that OSWAAL GSEB NCERT SOLUTIONS will help the students in school and after school in practicing and preparing extensively for both, Final Examinations as well as Competitive Examinations with utmost confidence! Some of the Key Highlights of Oswaal GSEB NCERT Solutions are:

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This book emphasises those features in solution chemistry which are

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difficult to measure, but essential for the understanding of both the qualitative and the quantitative aspects. Attention is paid to the mutual influences between solute and solvent, even at extremely small concentrations of the former. The described extension of the molecular concept leads to a broad view – not by a change in paradigm – but by finding the rules for the organizations both at the molecular and the supermolecular level of liquid and solid solutions.

Contents: Development and Present State
Atoms and Molecules
Chemical Bonding
Interactions between Molecules
The Liquid State
Anomalous Physical Properties of Liquid Water
Some Trivia about Water
The Phase Boundary of Liquid Water
Water in Biological Systems
Hydrophobic Solutes in Water
Hydrophilic Solutes in Water
Water and Alcohols
Characterization of Non-Aqueous Solvents
Solvation in Non-Aqueous Solvents
Ionization and Association in Non-Aqueous Solutions
Qualitative Aspects of the Molecular Concept
System Organization of Liquid Water
Changes in Organization of Liquid Water
Water within the Human Body
Organization in Non-Aqueous Solutions: Intramolecular System Organizations

Readership: Students and scientists in chemistry, physics, biology, pharmacy and medicine.

keywords: Solution Chemistry; Supermole; Liquid State; Hydrophobic Solutes; Hydrophilic Solutes; Ionization; Pharmacology; Liquid Properties; Solvents; Solvation

“Wherever possible, the authors have

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tried to make the text readable by using interesting illustrations to explain the relevance of the concepts that they describe ... this book will be excellent supplementary reading for undergraduates and will also be good preliminary background reading for researchers new to the area." Chemistry in Britian

Syllabus : Unit I : Solid State Unit II : Solutions Unit III : Electrochemistry Unit IV : Chemical Kinetics Unit V : Surface Chemistry Unit VI : General Principles and Processes of Isolation of Elements Unit VII : "p"-Block Elements Unit VIII : "d" and "f" Block Elements Unit IX : Coordination Compounds Unit X : Haloalkanes and Haloarenes Unit XI : Alcohols, Phenols and Ethers Unit XII : Aldehydes, Ketones and Carboxylic Acids Unit XIII : Organic Compounds Containing Nitrogen Unit XIV : Biomolecules Unit XV : Polymers Unit XV : Polymers Unit XVI : Chemistry in Everyday Life Content : 1. Solid State 2. Solutions 3. Electro-Chemistry 4. Chemical Kinetics 5. Surface Chemistry 6. General Principles And Processes Of Isolation Of Elements 7. P-Block Elements 8. D-And F-Block Elements 9. Coordination Compounds And Organometallics 10. Haloalkanes And Haloarenes 11. Alcohols, Phenols And Ethers 12. Aldehydes Ketones And

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Carboxylic Acids 13. Organic Compounds Containing Nitrogen 14.
Biomolecules 15. Polymers 16. Chemistry In Everyday Life Appendix :
1. Important Name Reactions And Process 2. Some Important Organic
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Electrochemistry Unit IV : Chemical Kinetics Unit V : Surface
Chemistry Unit VI : General Principles and Processes of Isolation of
Elements Unit VII : "p"-Block Elements Unit VIII : "d" and "f" Block
Elements Unit IX : Coordination Compounds Unit X : Haloalkanes and
Haloarenes Unit XI : Alcohols, Phenols and Ethers Unit XII :
Aldehydes, Ketones and Carboxylic Acids Unit XIII : Organic Compounds
Containing Nitrogen Unit XIV : Biomolecules Unit XV : Polymers Unit
XV : Polymers Unit XVI : Chemistry in Everyday Life Content : 1.
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Surface Chemistry 6. General Principles And Processes Of Isolation Of
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Coordination Compounds And Organometallics 10. Haloalkanes And
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